

SpellIndia

INDIA's **1**
No.

SPELLING BEE

Preparatory Study Material
Provider

www.phonicstore.com

ICSE ACADEMY
www.spellbeeacademy.com



MCQs

for

ICSE

Class 10

Chemistry

**MCQ
SET
2**

100+

Questions

**A Collection
of Questions
with Answers**



ICSE ACADEMY: How to Prepare for ICSE Class 10 exams

<https://www.spellbeeacademy.com/icse.html>



How to Prepare for ICSE Class 10 exams : Free Resources

Please click on subject to proceed further.

We will keep adding resources here till "March 2026".

So, save this link, keep visiting and stay updated.

(Resources include : Syllabus, Past Year Papers, Specimen Papers, Competency based Questions, Books pdf downloadable, 350+ Term Papers / Prelim Papers of various schools - across subjects, etc.)

- | | | | |
|--------------------------|-----------------------|--------------------------|-----------------------|
| 01 English Literature | 02 English Language | 03 Geography | 04 History & Civics |
| 05 Physics | 06 Chemistry | 07 Mathematics | 08 Biology |
| 09 Computer Applications | 10 Physical Education | 11 Hindi | 12 Commercial Studies |
| 13 Economics | 14 Technical Drawing | 15 Environmental Science | 16 Home Science |
| 17 Gujarati | 18 Marathi | 19 French | |



SCAN QR code to buy the book at amazon NOW.

SpellIndia
INDIA's No. 1 SPELLING BEE
Preparatory Study Material Provider
www.spellindia.com

ICSE ACADEMY
www.spellbeeacademy.com

Pati's

PREPARE
for
ICSE
Classes 9 & 10
ENGLISH GRAMMAR
(includes Board Specimen Papers of 5 years & Competency-focused questions)

28 YEARS Past Questions
75 Practice TESTS

Authoring to
solve ICSE exams
2024-2027
exams

Debashis Pati
Author is the 1st individual to write preparatory books on various topics of "multiple" Spelling Bee competitions in India. He has written the Maximum Number of Spelling Books as well as Tests in the world.

ing / Vocabulary / Grammar
id Olympiad Exam conductor.



SpellIndia
INDIA's No. 1 SPELLING BEE
Preparatory Study Material Provider
www.spellindia.com

ICSE ACADEMY
www.spellbeeacademy.com

Pati's

PREPARE
for
ICSE
Class 10
(Acts 3 to 5 only)
Julius Caesar

1000+ Practice QUESTIONS*
30 Practice TESTS
(From best traditions of 30 schools)

Past Years' Questions (13 years: 1990 onwards)
Competency focused Questions (1 year)
Multiple choice Questions (850+ nos*)
Extract based Questions (85+ extracts*)
*Excludes the questions in the 12 past years' questions and the 30 tests.

Authoring to
solve ICSE exams
2024-2027
exams

Debashis Pati
Author is the 1st individual to write preparatory books on various topics of "multiple" Spelling Bee competitions in India. He has written the Maximum Number of Spelling Books as well as Tests in the world.

ing / Vocabulary / Grammar
id Olympiad Exam conductor.



SpellIndia
INDIA's No. 1 SPELLING BEE
Preparatory Study Material Provider
www.spellindia.com

ICSE ACADEMY
www.spellbeeacademy.com

Pati's

PREPARE
for
ICSE
Class 10
HINDI GRAMMAR
350+ Sample practice questions & # 51 Tests

350+ SAMPLE QUESTIONS
51 TEST PAPERS

Authoring to
solve ICSE exams
2025 / 2026
onwards

तामसी पति
Tamasee Pati



SpellIndia
INDIA's No. 1 SPELLING BEE
Preparatory Study Material Provider
www.spellindia.com

ICSE ACADEMY
www.spellbeeacademy.com

Pati's

PREPARE
for
ICSE
Class 10
CIVICS

40 TEST PAPERS

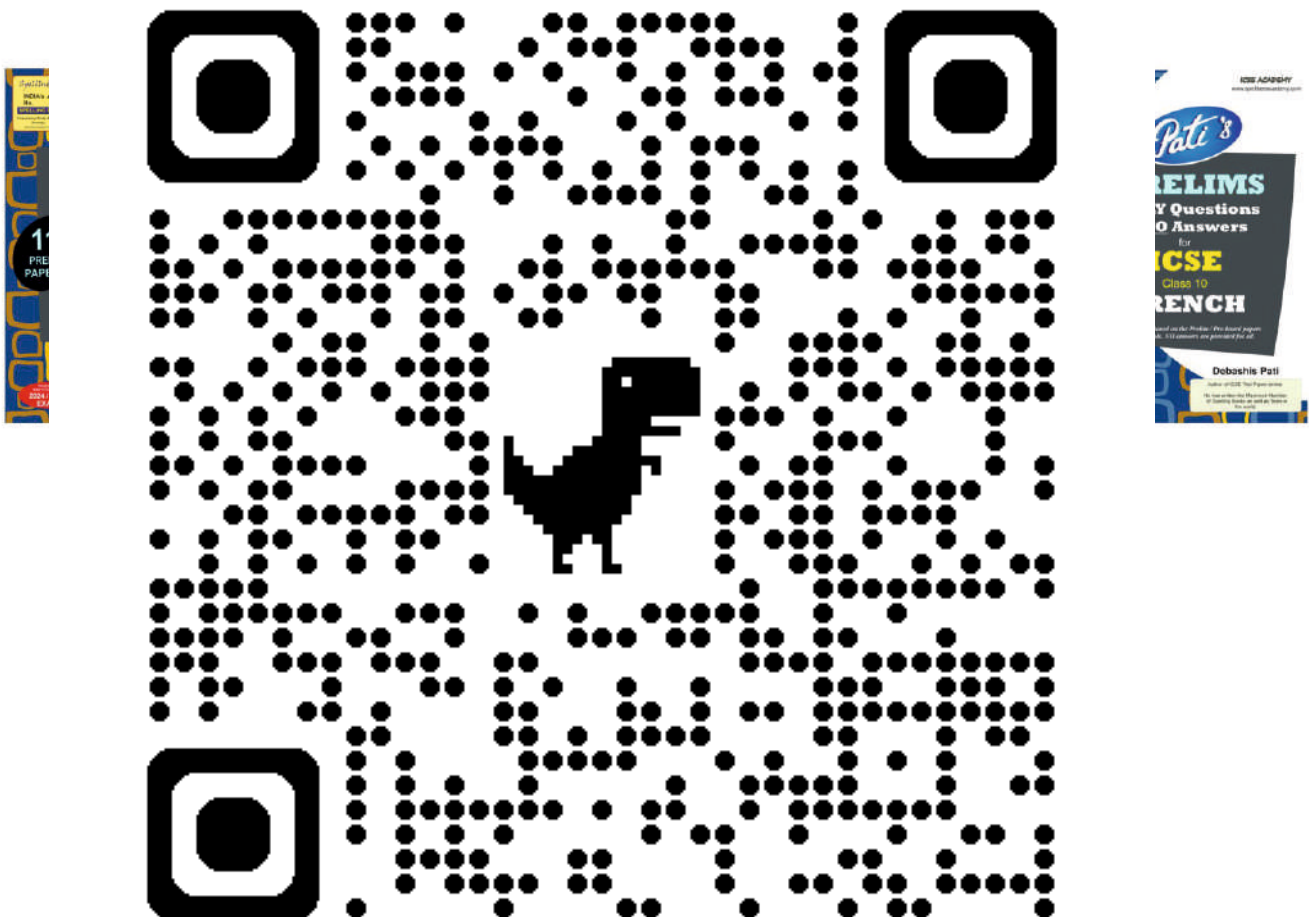
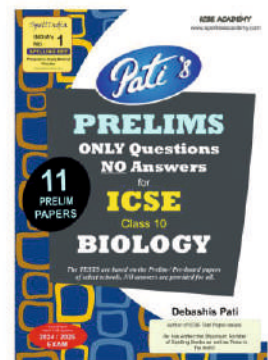
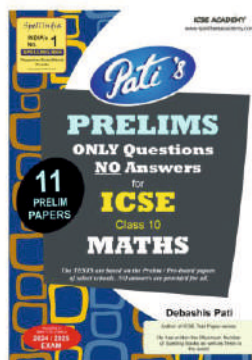
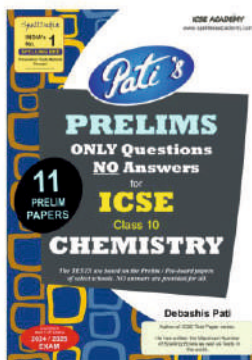
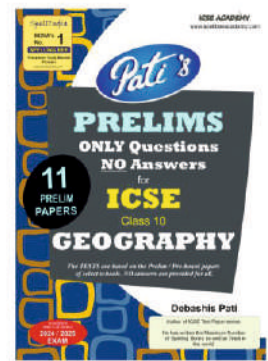
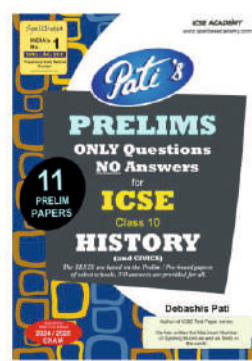
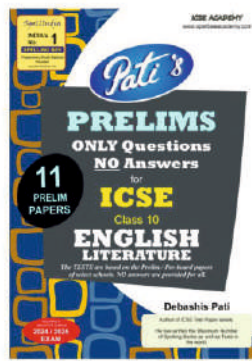
The TESTS are based on the Prelim / Pre-board papers of various schools. Answers are provided for all.
Competency Based Questions and 3 Specimen Papers are provided.

Authoring to
solve ICSE exams
2024 / 2025
EXAM

Debashis Pati
Author of ICSE Test Paper series.
n the Maximum Number
ooks as well as Tests in
the world.



Scan QR code for Free Access to 500+ Prelim Papers across 20 subjects



Prepare for ICSE CLASS 10 Chemistry

Free Resources

SCAN QR CODE Now





ICSE ACADEMY

Chemistry MCQs

MCQ
SET
2

100+ MCQs

Q. Choose the correct answers to the questions from the given options.

Note : Answers are provided at the start.

KARNATAKA ICSE SCHOOLS ASSOCIATION
ICSE STD. X Preparatory Examination 2026

Subject – CHEMISTRY (SCIENCE PAPER – 2)

Duration: 2 Hr.

Maximum Marks: 80

Date: 09.01.2026

General Instructions:

1. Answers to this paper must be written on the paper provided separately.
2. You will **not** be allowed to write during the first **15** minutes.
3. This time is to be spent in reading the Question Paper.
- 4. The time given at the head of this Paper is the time allotted for writing the answers.**
- 5. Section A is compulsory. Attempt any four questions from Section B.**
6. The intended marks for questions or parts of questions are given in brackets [].

SECTION A (40 Marks)

(Attempt all questions from this section)

Question 1

Choose the correct answers to the questions from the given options.

[15]

(Do not copy the question, write the correct answers only.)

- (i) Which of the following pairs are non-polar covalent compounds?
- | | |
|----------------------|-------------------------|
| 1. Hydrogen chloride | 2. Carbon tetrachloride |
| 3. Sodium chloride | 4. Methane |
- (a) 1 and 2
(b) 2 and 4
(c) 1 and 3
(d) 2 and 3
- (ii) Which of the following are the raw materials for the preparation of a fruit smelling compound in the presence a non-volatile strong acid?
- | | |
|----------------|------------|
| 1. Acetic acid | 2. Ethane |
| 3. Ethanol | 4. Ethanal |
- (a) 1 and 3
(b) 2 and 3
(c) 1 and 2
(d) 2 and 4

(iii) **Assertion (A)** Aluminium is extracted by electrolytic reduction method.

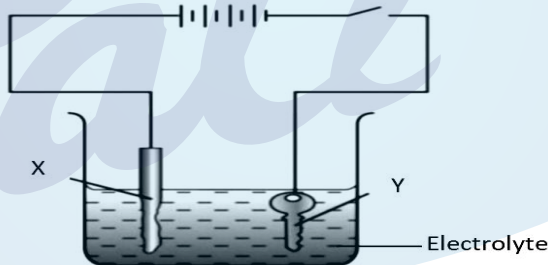
Reason (R) Aluminium is amphoteric in nature.

- (a) (A) is true and (R) is false.
- (b) (A) is false and (R) is true.
- (c) Both (A) and (R) are true and (R) is the correct explanation of (A).
- (d) Both (A) and (R) are true, but (R) is not the correct explanation of (A).

(iv) Identify the correct ionization equation for a divalent metal 'M'.

- (a) $M + 1e \longrightarrow M^{1+}$
- (b) $M - 2e \longrightarrow M^{2+}$
- (c) $M + 2e \longrightarrow M^{2+}$
- (d) $M - 2e \longrightarrow M^{2-}$

(v) A student Raghava wanted to electroplate locker key with gold. So he arranged the setup as shown in the below figure. Which of the following holds good for the same?



- (a) **Electrolyte**: Gold solution, **X** - Cathode: Impure gold metal rod, **Y** - Anode: Key
- (b) **Electrolyte**: Gold solution, **Y** - Cathode: Key, **X** - Anode: Pure gold metal rod
- (c) **Electrolyte**: Gold solution, **X** - Cathode: Key, **Y** - Anode: Impure gold metal rod
- (d) **Electrolyte**: Gold solution, **Y** - Cathode: Pure gold metal rod, **Y** - Anode: Key

(vi) The correct arrangement of elements with their no. of electron shells is:

- (a) He < C < Al < Ca
- (b) Ca < C < Si < P

(c) $\text{Na} > \text{F} > \text{Si} < \text{K}$

(d) $\text{H} < \text{Na} < \text{Li} < \text{N}$

(vii) Which of the following is correct in the electron dot diagram of acetate ion?

	C-H Bond	C-C Bond	C-O Single Bond	C-O Double Bond	Charge on acetate ion
(a)	3	1	1	1	+
(b)	3	2	2	1	+
(c)	2	1	1	2	-
(d)	3	1	1	1	-

(viii) The metal oxide which produces salt and water with acid as well as alkali is:

(a) K_2O

(b) Al_2O_3

(c) Na_2O

(d) Ca_2O

(ix) Which of the following pairs have the same empirical formula?

1. Methanol

2. Glucose

3. Acetic acid

4. Benzene

(a) 1 and 3

(b) 2 and 3

(c) 1 and 4

(d) 3 and 4

(x) Which of the following metal nitrate produces reddish brown coloured gas upon heating?

(a) Lead nitrate

(b) Potassium nitrate

(c) Sodium nitrate

(d) None of these

- (xi) Which of the following method is used in the concentration of bauxite ore?
- (a) Gravity Separation
 - (b) Froth Floatation
 - (c) Leaching
 - (d) Magnetic Separation
- (xii) The number of covalent bonds present in ammonium chloride is:
- (a) 3
 - (b) 4
 - (c) 5
 - (d) 2
- (xiii) The ion which do not form precipitate with weak alkali is:
- (a) Ferric ion
 - (b) Cuprous ion
 - (c) Calcium ion
 - (d) Ferrous ion
- (xiv) What property of dilute sulphuric acid can be demonstrated by treating it with ammonium carbonate salt to obtain carbon dioxide gas?
- (a) Oxidising property
 - (b) Acidic property
 - (c) Drying property
 - (d) Dehydrating property
- (xv) Dry ammonia gas does not change the red litmus to blue. Which of the following statements are true for this statement?

1. Ammonia gas is acidic in nature.
2. Ammonia gas does not ionise in gaseous state.
3. Ammonia gas is basic in nature.
4. Ammonia gas does not ionise in the absence of water.

- (a) Statements 1 and 2
- (b) Statements 1 and 4
- (c) Statements 3 and 2
- (d) Statements 3 and 4



KARNATAKA ICSE SCHOOLS ASSOCIATION
ICSE STD. X Preparatory Examination 2026
MARKING SCHEME – CHEMISTRY (SCIENCE PAPER 2)

Question 1		[15 X 1]
(i)	(b) 2 and 4	
(ii)	(a) 1 and 3	
(iii)	(d) Both (A) and (R) are true, but (R) is not the correct explanation of (A).	
(iv)	(b) $M - 2e \rightarrow M^{2+}$	
(v)	(b) Electrolyte : Gold solution, Y - Cathode: Key, X - Anode: Pure gold metal rod	
(vi)	(a) $He < C < Al < Ca$	
(vii)	(d) 3 1 1 1 -	
(viii)	(b) Al_2O_3	
(ix)	(b) 2 and 3	
(x)	(a) Lead nitrate	
(xi)	(c) Leaching	
(xii)	(a) 3	
(xiii)	(c) Calcium ion	
(xiv)	(b) Acidic property	
(xv)	(d) Statements 3 and 4	

KARNATAKA ICSE SCHOOLS ASSOCIATION
ICSE STD. X Preparatory Examination 2025
Subject- Chemistry

Duration 2 Hours

Maximum Marks:80

Date: 20.01.2025

Answer to this paper must be written on the paper provided separately.

*You will **not** be allowed to write during first 15 minutes.*

This time is to be spent in reading the question paper.

The time given at the head of this Paper is ~~not~~ allowed for writing the answer.

Section A is compulsory. Attempt **any four** questions from **Section B**.
The intended marks for questions or parts of questions given in the bracket [].

SECTION A (40 Marks)
(Attempt **all** questions from this **Section**).

Question 1

Choose the correct answers to the questions from the given options.

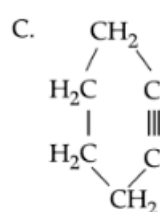
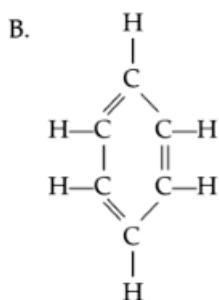
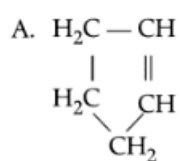
[15]

(Do not copy the question, write the correct answers only.)

- (i) During the electrolysis of copper sulphate solution using platinum electrodes, which of the following statements is correct?
P. When the deposition of Cu ions is completed then electrolysis of water takes place
Q. The blue colour of the solution gradually fades due to the reduction of Cu^{2+} ions
- (a) Only P
(b) Only Q
(c) Both P and Q
(d) Neither P and Q

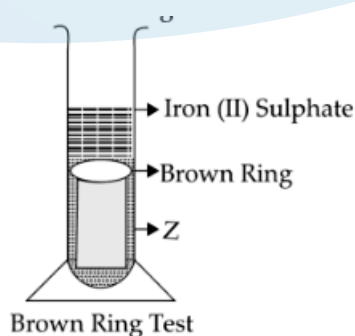
- (ii) Arun tested a salt solution by adding ammonium hydroxide dropwise, forming a dirty green precipitate. On adding excess ammonium hydroxide, the precipitate is insoluble. The cation present in the salt is:
- (a) Zinc ion
 - (b) Copper(II) ion
 - (c) Calcium ion
 - (d) Iron(II) ion
- (iii) Calculate the volume occupied by 3.4 g of ammonia at S.T.P.
- (a) 4.48 litre
 - (b) 6.05 litre
 - (c) 5.48 litre
 - (d) 3.63 litre
- (iv) The equation below shows the reaction between element 'M' and dilute hydrochloric acid.
- $$M(s) + 2HCl(aq.) \rightarrow MCl_2(aq.) + H_2(g)$$
- Which particles are responsible for conducting electricity in dilute hydrochloric acid and compound MCl_2 ?
- (a) Electrons
 - (b) Both positive and negative ions
 - (c) Only negative ions
 - (d) Only positive ions
- (v) **Assertion (A)** Hydrogen chloride gas is purified by passing it through conc. H_2SO_4 .
Reason (R) Concentrated sulphuric acid is used as the drying agent.
- (a) Both A and R are true and R is the correct explanation of A
 - (b) Both A and R are true and R is not the correct explanation of A
 - (c) A is true but R is false
 - (d) A is false but R is true
- (vi) An element A belonging in Period 3 and Group II will have:
- (a) 2 shells and 3 valence electrons
 - (b) 3 shells and 3 valence electrons
 - (c) 2 shells and 2 valence electrons
 - (d) 3 shells and 2 valence electrons

(vii) Which of the following organic compounds is of benzene?



- (a) Compound A
 (b) Compound B
 (c) Compound C
 (d) Compound B and C
- (viii) Which of the following is the atomic number of an element that forms basic oxide?
- (a) 18
 (b) 17
 (c) 19
 (d) 15
- (ix) A sulphate which forms a gelatinous white precipitate with sodium hydroxide and is also soluble in excess of it:
- (a) Ferrous sulphate
 (b) Zinc sulphate
 (c) Copper sulphate
 (d) Ferric sulphate
- (x) Which of the following electronic configuration suggests an element with the highest electropositivity ?
- (a) 2,1
 (b) 2,2
 (c) 2,7
 (d) 2,8

- (xi) **Assertion (A)** Alkali metals are stored under oil to prevent reaction with air.
Reason (R) Alkali metals slowly react with oxygen and moisture in the air.
- (a) Both A and R are true and R is the correct explanation of A
 (b) Both A and R are true and R is not the correct explanation of A
 (c) A is true but R is false
 (d) A is false but R is true
- (xii) What is the ratio between the volumes occupied by carbon dioxide (CO_2) and hydrogen (H_2) when 22 grams of CO_2 and 4 grams of H_2 are measured under the same conditions of temperature and pressure?
- (a) 4:1
 (b) 1:4
 (c) 2:16
 (d) 16:2
- (xiii) Aqueous lead (II) nitrate can be distinguished from aqueous zinc nitrate by adding any of the following solution in excess, except:
- (a) Aqueous sodium sulphate
 (b) Dilute sulphuric acid
 (c) Sodium hydroxide solution
 (d) Aqueous potassium chloride
- (xiv) Which ion is determined by the test shown in the image?



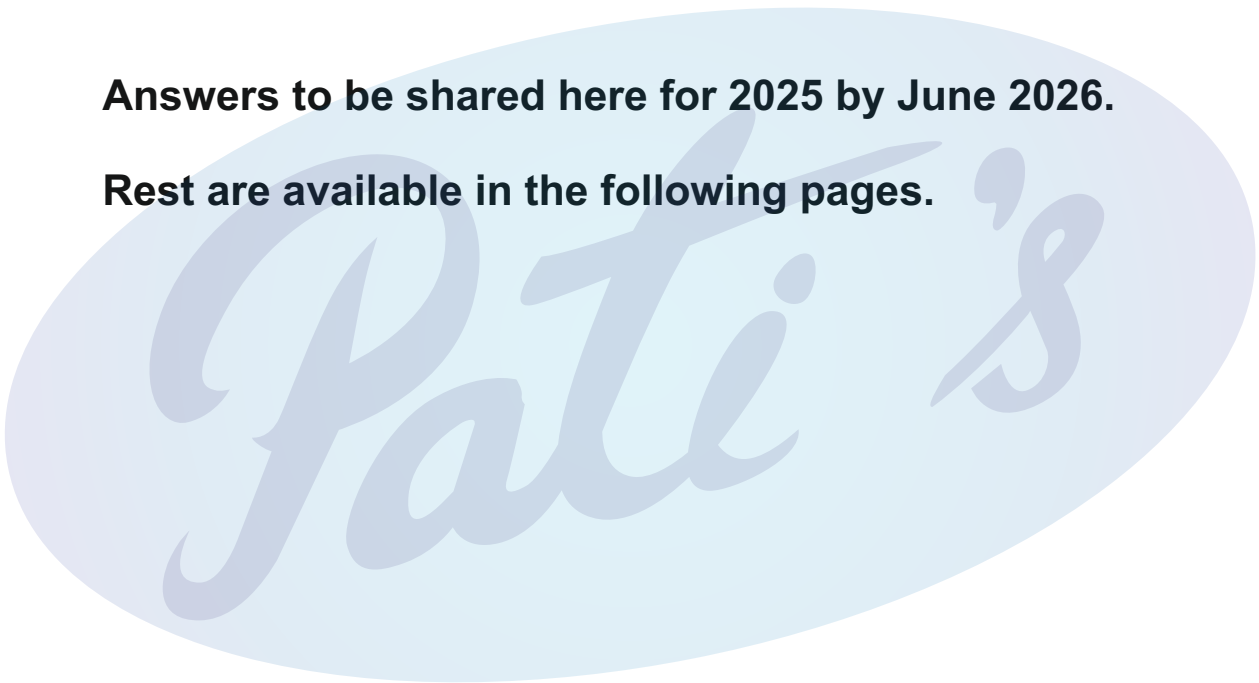
- (a) Chloride
 (b) Nitrate
 (c) Sulphate
 (d) Phosphate

- (xv) What distinguish amphoteric oxides from other types of oxides?
- (a) They react with acids to form salt and water.
 - (b) They react with bases to form salt and water.
 - (c) They react with both acids and bases to form salt and water.
 - (d) They do not react with either acids or bases.



Answers to be shared here for 2025 by June 2026.

Rest are available in the following pages.



KARNATAKA ICSE SCHOOLS ASSOCIATION

ICSE STD. X Preparatory Examination 2024

Subject: CHEMISTRY (SCIENCE PAPER 2)

Maximum Marks: 80

Time Allowed: Two hours

Date: _____

General Instructions:

Answers to this Paper must be written on the paper provided separately.

You will not be allowed to write during first 15 minutes.

This time is to be spent in reading the question paper.

The time given at the head of this Paper is the time allowed for writing the answers.

Section A is compulsory. Attempt **any four** questions from **Section B**.

The intended marks for questions or parts of questions are given in brackets [].

SECTION A

(Attempt **all** questions from this Section.)

Question 1

Choose the correct answers to the questions from the given options. [15]

(Do not copy the question, write the correct answers only.)

- i. Electrolysis of acidified water produces Oxygen gas at the anode. Which of the following ions move towards the anode?
 - P. SO_4^{2-}
 - Q. OH^{1-}
 - R. O^{2-}
 - (a) Only R
 - (b) Both P and R
 - (c) Both P and Q
 - (d) Only Q
- ii. A compound X is heated in a dry test tube. The gas liberated is passed into another test tube containing potassium iodide solution that produces violet vapours. The liberated gas could possibly be?
 - (a) Oxygen
 - (b) Nitrogen dioxide
 - (c) Chlorine
 - (d) Sulphur dioxide

iii. The atomic masses of oxygen(O), fluorine(F) and neon(Ne) are 16, 19 and 20 respectively.

Which of the following statements regarding the volume of gases at STP in 16 g of oxygen, 19 g of fluorine and 20 g of neon is correct?

M. 16 g of oxygen will occupy half the volume occupied by 20 g of neon.

N. 20 g of neon will occupy twice the volume occupied by 19 g of fluorine.

- (a) Both M and N
- (b) Only M
- (c) Only N
- (d) Neither M nor N

iv. Hydrogen chloride gas is passed through conc. sulphuric acid and collected in a jar. Moist blue litmus paper is exposed to the gas in the jar. What do you observe?

- (a) It remains blue.
- (b) It gets bleached
- (c) It turns red
- (d) It turns red then gets bleached

v. Sulphur reacts with hot conc. nitric acid producing sulphuric acid, water and nitrogen dioxide. The property of nitric acid emphasised here is

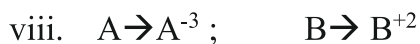
- (a) Acid property
- (b) Oxidizing nature
- (c) Reducing nature
- (d) Dehydrating property

vi. Which among the following conditions does not hold true for electroplating?

- (a) Direct current should be used.
- (b) The electrolyte must contain ions of the base metal to be electroplated.
- (c) A low current for a prolonged period of time must be used.
- (d) The metal to be plated on the article is made the anode.

vii. The acidity is 3 for

- (a) Ammonium hydroxide
- (b) Copper hydroxide
- (c) Zinc hydroxide
- (d) Ferric hydroxide



Number of electron present in the valence shell and the type of element of atoms A and B respectively are:

	Number of electrons in valence shell	Type of Elements
W	3,2	Non-metal, non-metal
X	5,2	Non-metal, metal
Y	3,6	Metal, non-metal
Z	5,3	Non-metal, metal

- (a) W
- (b) X
- (c) Y
- (d) Z

ix. Salt of Fe^{2+} ion will be of the colour

- (a) Green
- (b) Brown
- (c) Pink
- (d) Yellow

x. An element with atomic number ____ will form a base when its oxide is dissolved in water.

- (a) 17
- (b) 7
- (c) 15
- (d) 19

xi. Reaction between ammonia and chlorine produces a yellow explosive liquid of nitrogen trichloride. The ratio of ammonia: chlorine is

- (a) 1:2
- (b) 2:1
- (c) 1:3
- (d) 3:1

- xii. -CHO functional group is present in
- Butanal
 - Ethanol
 - Propanone
 - Methanoic acid
- xiii. $\text{Cu}^{2+} \rightarrow \text{Cu}$. This is an example for
- Oxidation
 - Reduction
 - Redox reaction
 - None of the above
- xiv. The promoter used in Haber's process
- Pt
 - Fe
 - K_2O
 - Mo
- xv. A negative divalent element will belong to which group?
- Group 2
 - Group 13
 - Group 14
 - Group 16

KARNATAKA ICSE SCHOOLS ASSOCIATION

ICSE STD. X Preparatory Examination 2024

Subject: CHEMISTRY (SCIENCE PAPER 2)

Maximum Marks: 80

Time Allowed: Two hours

Date: _____

ANSWER KEY

SECTION A

(Attempt **all** questions from this Section.)

Question 1

Choose the correct answers to the questions from the given options.

[15]

(Do not copy the question, write the correct answers only.)

i.	(d) Only Q
ii.	(b) Nitrogen dioxide
iii.	(a) Both M and N
iv.	(c) It turns red
v.	(b) Oxidizing nature
vi.	(b) the electrolyte must contain ions of the base metal to be electroplated.
vii.	(d) Ferric hydroxide
viii.	(b) X
ix.	(a) Green
x.	(d) 19
xi.	(c) 1:3
xii.	(a) Butanal
xiii.	(b) Reduction
xiv.	(d) Mo
xv.	(d) Group 16

General Instructions:

Answers to this Paper must be written on the paper provided separately.

You will not be allowed to write during first 15 minutes.

This time is to be spent in reading the question paper.

The time given at the head of this Paper is the time allowed for writing the answers.

Section A is compulsory. Attempt any four questions from Section B.

The intended marks for questions or parts of questions are given in brackets [].

SECTION A

(Attempt all questions from this Section.)

Question 1

Choose the correct answers to the questions from the given options.

[15]

(Do not copy the question, write the correct answers only.)

i. Electrolysis of acidified water produces Oxygen gas at the anode.

Which of the following ions move towards the anode?

P. SO_4^{2-}

Q. OH^-

R. O^{2-}

(a) Only R

(b) Both P and R

(c) Both P and Q

(d) Only Q

ii. A compound X is heated in a dry test tube. The gas liberated is passed into another test tube containing potassium iodide solution that produces violet vapours. The liberated gas could possibly be?

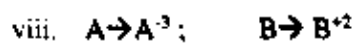
(a) Oxygen

(b) Nitrogen dioxide

(c) Chlorine

(d) Sulphur dioxide

- iii. The atomic masses of oxygen(O), fluorine(F) and neon(Ne) are 16, 19 and 20 respectively.
Which of the following statements regarding the volume of gases at STP in 16 g of oxygen, 19 g of fluorine and 20 g of neon is correct?
M. 16 g of oxygen will occupy half the volume occupied by 20 g of neon.
N. 20 g of neon will occupy twice the volume occupied by 19 g of fluorine.
- (a) Both M and N
 - (b) Only M
 - (c) Only N
 - (d) Neither M nor N
- iv. Hydrogen chloride gas is passed through conc. sulphuric acid and collected in a jar. Moist blue litmus paper is exposed to the gas in the jar. What do you observe?
- (a) It remains blue.
 - (b) It gets bleached
 - (c) It turns red
 - (d) It turns red then gets bleached
- v. Sulphur reacts with hot conc. nitric acid producing sulphuric acid, water and nitrogen dioxide. The property of nitric acid emphasised here is
- (a) Acid property
 - (b) Oxidizing nature
 - (c) Reducing nature
 - (d) Dehydrating property
- vi. Which among the following conditions does not hold true for electroplating?
- (a) Direct current should be used.
 - (b) The electrolyte must contain ions of the base metal to be electroplated.
 - (c) A low current for a prolonged period of time must be used.
 - (d) The metal to be plated on the article is made the anode.
- vii. The acidity is 3 for
- (a) Ammonium hydroxide
 - (b) Copper hydroxide
 - (c) Zinc hydroxide
 - (d) Ferric hydroxide



Number of electron present in the valence shell and the type of element of atoms A and B respectively are:

	Number of electrons in valence shell	Type of Elements
W	3,2	Non-metal, non-metal
X	5,2	Non-metal, metal
Y	3,6	Metal, non-metal
Z	5,3	Non-metal, metal

- (a) W
(b) X
(c) Y
(d) Z
- ix. Salt of Fe^{2+} ion will be of the colour
(a) Green
(b) Brown
(c) Pink
(d) Yellow
- x. An element with atomic number _____ will form a base when its oxide is dissolved in water.
(a) 17
(b) 7
(c) 15
(d) 19
- xi. Reaction between ammonia and chlorine produces a yellow explosive liquid of nitrogen trichloride. The ratio of ammonia: chlorine is
(a) 1:2
(b) 2:1
(c) 1:3
(d) 3:1

- xii. -CHO functional group is present in
- (a) Butanal
 - (b) Ethanol
 - (c) Propanone
 - (d) Methanoic acid
- xiii. $\text{Cu}^{2+} \rightarrow \text{Cu}$. This is an example for
- (a) Oxidation
 - (b) Reduction
 - (c) Redox reaction
 - (d) None of the above
- xiv. The promoter used in Haber's process
- (a) Pt
 - (b) Fe
 - (c) K_2O
 - (d) Mo
- xv. A negative divalent element will belong to which group?
- (a) Group 2
 - (b) Group 13
 - (c) Group 14
 - (d) Group 16

Answers

(I) d (ii) b (iii) a (iv) c (v) b (vi) b (vii) d (viii) b
(ix) a (x) d (xi) c (xii) a (xiii) b (xiv) d (xv) d

Section A

Question 1: Choose one correct answer to the questions from the given options. [15]

i) In Electro refining of copper using acidified copper, which of the following is /are correct?

- (I) Pure copper is used at cathode (II) Impure copper is used as anode .
 (III) Reduction takes place at cathode. (IV) Platinum electrode is used as anode.
 (a) I,II, and III (b) II,III, and IV (c) I,III,IV (d) I,II,IV

ii) Some students were working in the lab, the teacher gave them a compound 'P' and asked them to add conc. Sulphuric acid , then she asked them to test the gas formed using moist blue litmus paper. The students observed that the blue litmus changed its color to red. Help the students to identify the compound 'P':

- (a) Zinc sulphate (b) Ammonium Carbonate (c) Copper nitrate (d) Sodium chloride

iii) Ammonia gas is passed through following solution. Which of the following observation is correct?

	Red litmus solution	Phenolphthalein solution
X	Blue	Pink
Y	Red	Colorless
Z	Blue	Blue
M	Red	Green

- a) X b) Y c) Z d) M

iv) The solution in which Aluminium ore dissolves

- (a) Calcium hydroxide (b) Sodium hydroxide (c) Copper sulphate (d) Silver nitrate

v) An electrolyte which dissociated completely in to ions :

- (a) Acetic acid (b) Ammonium hydroxide (c) Sodium hydroxide (d) Oxalic acid

vi) An element 'X; attracts the electron towards itself, it must belong to :

- (a) Group 1 (b) Group 17 (c) Group 2 (d) Group 18

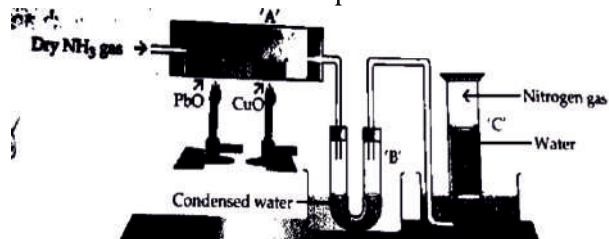
vii) The molecular formula of an organic compound is $C_{49}H_{106}$, the hydrocarbon will belong to :

- (a) Alkanes (b) Alkenes (c) Alkynes (d) None of the above

viii) The molecular mass of a compound is 88 amu having empirical formula C_2H_4O . Its molecular formula would be.

- (a) $C_2H_4O_2$ (b) $C_4H_8O_2$ (c) $C_8H_4O_2$ (d) $C_2H_8O_2$

(ix) The role of CuO or PbO in the experiment shown in diagram is :



a) Oxidizing agent (b) Dehydrating agent c) Reducing agent d) Catalyst

x) The product formed at the cathode in electroplating of an article with nickel is :

(a) Hydrogen gas (b) Nickel ions (c) Nickel atoms (d) Oxygen gas

xi) Ammonia molecule contains ____

- (a) A triple covalent bond and one lone pair of electrons
 (b) Three single covalent bonds and two lone pair of electrons
 (c) Three single covalent bonds and one lone pair of electrons
 (d) A double covalent bond and one lone pair of electrons

xii) The method of preparation used for preparation of iron(II) chloride is :

- (a) Synthesis reaction (b) Displacement reaction
 (c) Double displacement reaction (d) Decomposition reaction.

xiii) During the electrolysis of acidified water which of the following takes place:

- (a) Oxygen is released at cathode (b) Oxygen is released at anode
 (c) Sulphur dioxide is released at anode. (d) Hydrogen is released at anode.

xiv) Carbon to carbon double bond is found in:

- (a) 2-butylene (b) Acetic acid (c) Acetaldehyde (d) Ethyl alcohol

xv) Assertion:(A) Few drops of dilute acid is added to a solution of zinc sulphide , a colorless gas is formed with rotten egg odour.

Reason (R) Gas formed does not turn moist lead acetate paper silvery black.

- (a) Both A and R are true
 (b) A and R are true ,but R is the correct explanation of A
 (c) A is true , but R is not the correct explanation of A
 (d) Both A and R are false

Answers

(I) a (ii) d (iii) a (iv) b (v) c (vi) b (vii) d (viii) b
(ix) a (x) c (xi) c (xii) b (xiii) b (xiv) a (xv) c

Section A(40 Marks) [Attempt *all* questions from this section]

Question 1: Choose one correct answer to the questions from the given options. [15]

- (i) Solid lead (II) bromide cannot be electrolysed but molten lead(II) bromide can be electrolysed. What is the reason for this difference?
- (a) Solid lead(II) bromide does not contain ions
 (b) The ions can move in molten lead(II) bromide but not in the solid
 (c) All the ionic bonds are broken in molten lead (II) bromide
 (d) Electrodes can be put into molten lead (II) bromide but not into solid
- (ii) Unsaturated hydrocarbons under go:
- (a) Addition (b) Substitution (c) Reduction (d) Decomposition
- (iii) An object is electroplated with copper which statement about the process is correct.
- (a) The concentration of the copper ions in the electrolyte decreases.
 (b) Copper is used as a cathode
 (c) Copper ions gain electrons from the object being electroplated
 (d) Oxygen is evolved at the anode
- (iv) The impurity which separates out on addition of conc. solution of alkali to impure bauxite is
- (a) PbO (b) Fe₂O₃ (c) Al₂O₃ (d) ZnO
- (v) The metal hydroxide that reacts with both acids and alkalis to form salt and water is:
- (a) Calcium hydroxide (b) Magnesium hydroxide
 (c) Aluminium hydroxide (d) Ferric hydroxide
- (vi) Fuming nitric acid renders _____ passive.
- (a) Na and K (b) Mg and Mn (c) Cu and Zn (d) Fe and Al
- (vii) The product obtained in the absorption tower in contact process is: _____
- (a) Pyro sulphuric acid (b) Sulphuric acid
 (c) Sulphur trioxide (d) Sulphur dioxide
- (viii) Element with atomic number 11 _____
- (a) Will accept electron and get reduced (b) Will lose an electron and get oxidised
 (c) Will accept an electron and get oxidised (d) Will lose an electron and get reduced.
- (ix) Aqueous solution of cupric chloride forms a deep inky blue solution on addition of:
- (a) Drop wise sodium hydroxide (b) Excess sodium hydroxide
 (c) Dropwise ammonium hydroxide (d) Excess ammonium hydroxide
- (x) Assertion: Nitric acid reduces hydrogen to form water
 Reason: Nitric acid decomposes easily to form nascent oxygen.
- (a) Both 'A' and 'R' are false (b) Both 'A' and 'R' are correct
 (c) 'A' is correct and 'R' is false (d) 'A' is false and 'R' is correct
- (xi) Compound 'Q' on thermal decomposition releases a neutral gas as the only gas, compound 'Q' could be:
- (a) Copper nitrate (b) Sodium nitrate (c) Calcium nitrate (d) Zinc nitrate
- (xii) Assertion: All covalent compounds dissolve in water
 Reason: The molecules in them have very weak Vander Waal's force of attraction

- (a) Both A and R are correct. (b) A is false R is correct
(c) Both A and R are false (d) A is true and R is false
- (xiii) The molecular mass of a compound is 88 amu, having empirical formula C_2H_4O . Its molecular formula would be. [C = 12, H = 1, O = 16]
(a) $C_2H_4O_2$ (b) $C_4H_8O_2$ (c) $C_8H_4O_2$ (d) $C_2H_8O_2$
- (xiv) Which salt can be prepared by an acid and alkali titration method?
(a) Ammonium sulphate (b) Copper (II) Sulphate
(c) Iron (II) Sulphate (d) Zinc Sulphate
- (xv) Which statement applies to all three of compounds ethane, ethene and ethanol?
(a) One molecule of each compound contains the same number of carbon atoms.
(b) One mole of each compound contains the same number of hydrogen atoms.
(c) They all occur in crude oil.
(d) They all are liquid at room temperature.



Answers

(I) a (ii) a (iii) c (iv) b (v) c (vi) d (vii) a (viii) b
(ix) d (x) d (xi) b (xii) b (xiii) b (xiv) a (xv) a

12. Three chemicals P, Q & R were each dissolved in water. The table shows some of the reactions of these solutions

Solution	Reaction when sodium carbonate is added	Reaction with solid ammonium chloride
P	Gas evolved	No reaction
Q	No reaction	Gas evolved
R	No reaction	No reaction

The pH of the three solutions was also measured. What are the correct pH values of these solutions?

- A. P =2, Q =7, R=13 B. P=2, Q=13, R=7 C. P=7. Q=2, R=13 D. P=13. Q=7. R=2

13. When cold dilute nitric acid is added to a solid, S, a flammable gas, G, is formed. Gas G is less dense than air. What are S and G

	Solid S	Gas G
A	Copper	Hydrogen
B	Copper carbonate	Carbon dioxide
C	Magnesium	Hydrogen
D	Zinc carbonate	Carbon dioxide

14. The gas which has rotten egg smell

- A. Hydrogen sulphide B. Hydrogen Chloride C. Hydrogen fluoride D. Hydrogen Bromide

15. The froth floatation method is generally used for concentration of the ___ ores.

- A. Carbonate B. Sulphide C. Sulphate D. Nitrate

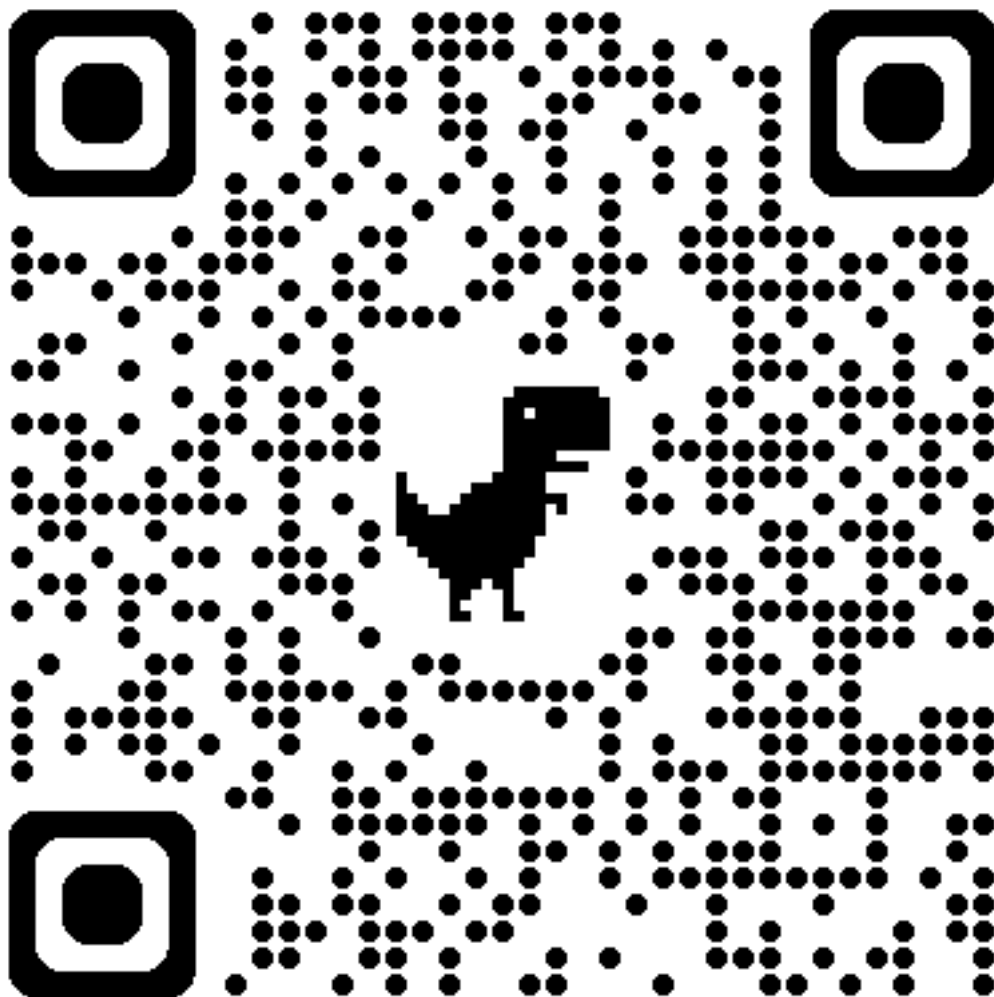
Answers

1 A 2 B 3 A 4 D 5 B 6 A 7 A 8 C 9 C 10 D
11 C 12 B 13 C 14 A 15 B

Prepare for ICSE CLASS 10 Chemistry

Free Resources

SCAN QR CODE Now



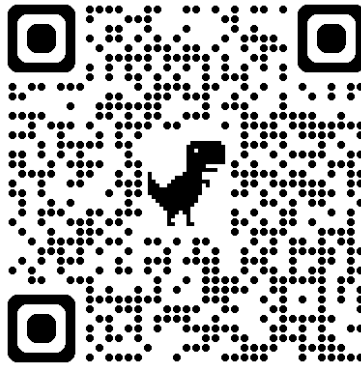


Prepare for ICSE CLASS 10
Free Resources

SCAN QR CODE Now



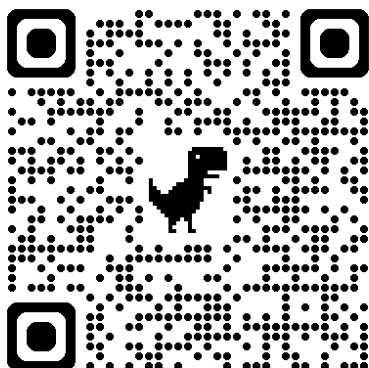
History/Civics



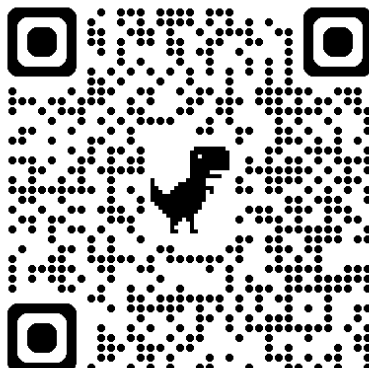
Geography



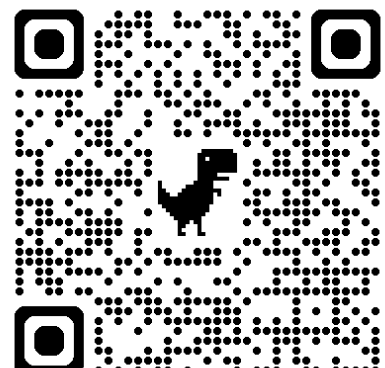
Maths



Physics



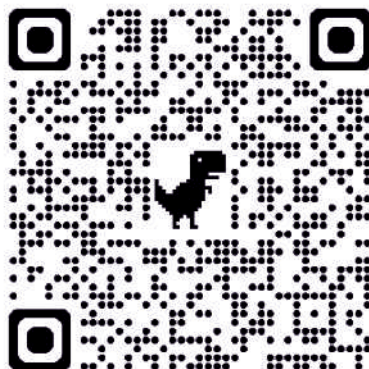
Chemistry



Biology



Hindi



Physical
Education



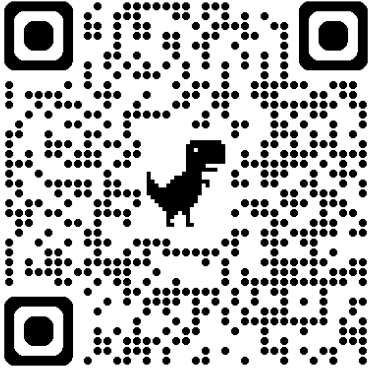
Computer
Applications





Prepare for ICSE CLASS 10
Free Resources

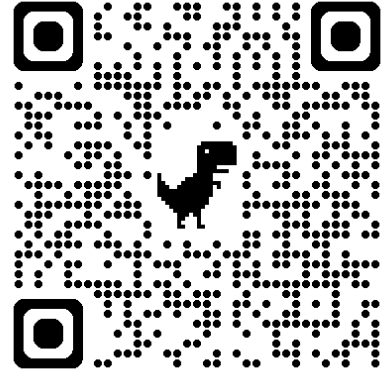
SCAN QR CODE Now



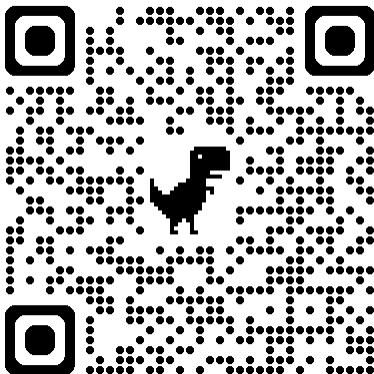
Economics



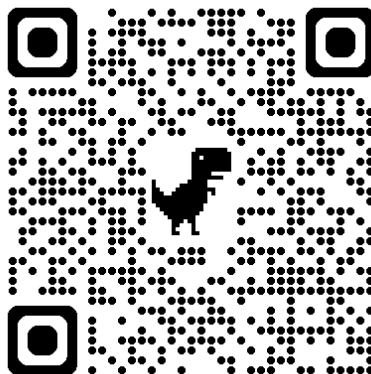
Commercial
Studies



French



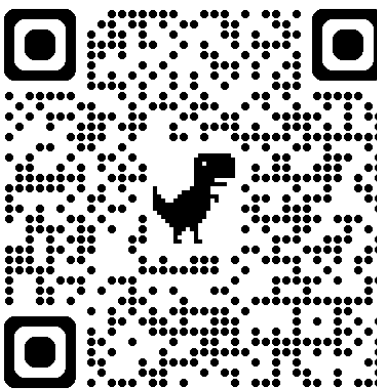
Robotics & AI



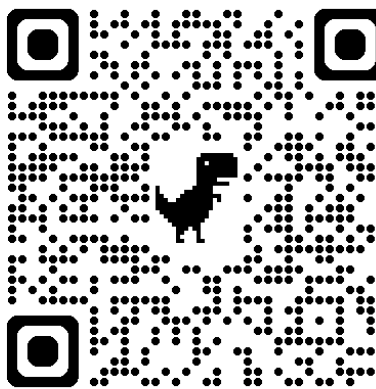
Home Science



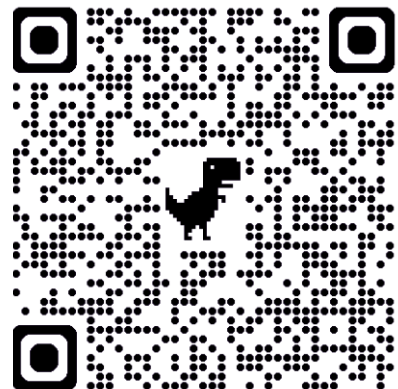
EVS



Marathi



Gujarati



Odiya



Scan QR code for Free Access to 500+ Prelim Papers across 20 subjects

